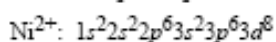
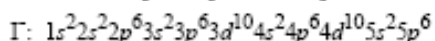
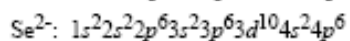
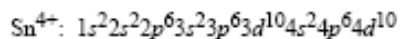
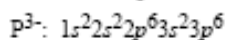
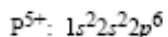


Answers to selected suggested problems in Chapter 2 (1, 2, 5, 7, 8, 9, 18, 19, 23)

Note: The key to some answers are not given since they come from the book; they should be apparent from the chapter text or solutions given in the back of book.

2.7



2.8

- a. Ar
- b. Xe

2.9

Each element has two s electrons.

2.19

- a. 73.4%
- b. 6.1%
- c. 93.4%
- d. 14.8%
- e. 51.4%

2.23

The intermolecular bonding for HF is hydrogen bonding. For HCl, the intermolecular bonding is van der Waals. Since hydrogen bonding is stronger than van der Waals, HF will have a higher melting temperature.